

Chapter 7. Atomic Structure

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Solution 1:

S.N.	Properties	Electron	Proton	Neutron
1.	Position	Electrons revolve around the nucleus in orbits.	A proton resides in the central part of nucleus.	Neutron reside in nucleus.
2.	Charge	An electron possesses a unit negative charge of magnitude - 1.602×10^{-19} coulombs.	A proton possesses a unit positive charge of the value 1.602×10^{-19} coulombs.	A neutron possesses zero charge. So, it is electrically neutral particle.
3.	Mass	Electron has negligible mass. The mass of an electron is $1/1837$ times the mass of a hydrogen atom i.e. 9.108×10^{-28} g.	Mass of proton is nearly same as that of hydrogen atom i.e. 1 amu or 1.672×10^{-24} g	Mass of neutron is slightly greater than proton i.e. 1.675×10^{-24} g.
4.	Notation	An electron is indicated as ${}^0_{-1}\text{e}$ where zero indicates that it has no mass and -1 indicates that it has one unit of negative charge.	A proton is indicated by ${}^{+1}_1\text{p}$ or ${}^{+1}_1\text{H}$ where 1 indicates that it has a mass equal to 1 amu and +1 stands for one unit positive charge.	A neutron is indicated by ${}^1_0\text{n}$ where 0 stands for zero charge and 1 stands for 1 amu mass.

Solution 2:

Number of electrons=12

Number of protons=12

Number of neutrons=24-12=12



Number of electrons=19

Number of protons=19

Number of neutrons=39-19=20



Number of electrons=20

Number of protons=20

Number of neutrons=40-20=20



Number of electrons=15

Number of protons=15

Number of neutrons=31-15=16

Solution 3:

1. **(a) Mass of an atom** – Protons and neutrons
2. **(b) Size of an atom** – Electrons

Solution 4:

Three fundamental particles of an atom are-

1. Proton
2. Electron
3. Neutron

Solution 5:

1. **(a) An atom** – An atom is the smallest particle of an element which may or may not be capable of independent existence.
2. **(b) An element** – An element is usually defined as a pure substance that contains only one kind of particles. These particles may be atoms or molecules.

Solution 6:

Atomic number – The number of protons present in the nucleus of an atom is called the atomic number of the atom.

It is denoted by 'Z'.

Solution 7:

The protons and neutrons are collectively called as nucleons.

Solution 8:

Isotopes are atoms of the same element, having the same atomic number, same chemical properties but different mass number, i.e., the atoms differ in the number of neutrons. The three isotopes of hydrogen atoms are-

1. Protium
2. Deuterium
3. Tritium

Solution 9:

Electrons take part in a chemical reaction. Therefore, the chemical properties of an element depend upon the electronic configuration. Since, isotopes of an element have the same atomic number and hence same electronic configuration. So, they exhibit the same chemical properties.

Solution 10:

S.N.	Dalton's atomic theory	Modern Atomic Theory
1.	According to Dalton's atomic theory, all matter is composed of extremely small particles called atoms.	According to modern theory, an element is made up of extremely small particles called atoms.
2.	Atoms of a given element are identical, both in mass and properties.	All atoms of an element may not be alike; they can have different masses (i.e. Isotopes)
3.	Atoms are indivisible and indestructible.	Atoms are divisible and are no more indestructible. Atom consists of still smaller particles like proton, electron and neutron.
4.	Atom is the smallest particle of an element.	Atoms take part in all chemical reactions, the relative numbers and kinds of atoms are constant in a given compound.
5.	Atoms take part in all chemical reactions, the relative no. and kinds of atoms are constant in a given compound.	

Solution 11:

The outermost orbit of an element is called valence shell. The electrons present in the outermost orbit (valence shell) of an element are called valence electrons.

Solution 12:

Atoms may have incomplete octet. During the formation of a molecule, an atom of a particular element gains, loses or shares electrons until it acquires a stable configuration of eight electrons in its valence shell.

Solution 13:

1. **(a)** Atomic number = Number of protons = 20
2. **(b)** Mass number = Number of protons + Number of electrons = 20 + 20 = 40
3. **(c)** Electronic configuration = 2, 8, 8, 2
4. **(d)** Valency = 2

Solution 14:

1. **(a)** A=1
B=3
C=2

$$D=1$$

$$E=1$$

2. **(b)** E contains the greatest number of neutrons.
3. **(c)** A contains the least number of electrons.
4. **(d)** B contains equal number of electrons and neutrons.
5. **(e)** A is a metal.
6. **(f)** Fluorine is the most reactive of the non-metals.
7. **(g)** Both are electronegative and non-metals.
8. **(h)** D and E belongs to a particular family.
9. **(i)** A_3B will be the formula if atoms of A combines with atoms of B.
10. **(j)** 'B' will combine in a trivalent element while 'C' will combine as a bivalent element.

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Solution 15:

"During the formation of the molecule, an atom of a particular element gains, loses electrons or shares electrons until it acquires a stable configuration of eight electrons in its valence shell" i.e. until it acquires octet.

Solution 16:

K can accommodate maximum of 2 electrons.

L can accommodate maximum of 8 electrons.

M can accommodate maximum of 18 electrons.

Solution 17:

Cathode rays are formed at the negative electrode of the discharge tube experiment.

Solution 18:

1. In K maximum number of 2 electrons can be accommodated.
2. In L maximum number of 8 electrons can be accommodated.
3. In M maximum number of 18 electrons can be accommodated.
4. In N maximum number of 32 electrons can be accommodated.

Solution 19:

Suppose the percentage of $^{10}_5B$ be A and the percentage of $^{11}_5B$ be (100-A). It implies that,

$$\frac{A \times 10 + (100-A) \times 11}{100} = 10.8$$

$$11A - 10A = 1100 - 1080$$

$$A = 20$$

So, the percentage of $^{10}_5B = 20\%$

The percentage of $^{11}_5B = 100 - 20 = 80\%$

Solution 20:

Electronic configuration of magnesium is -2, 8, 2. Since, it has 2 electrons in its valence shell, so its valency is 2.

Solution 21:

Number of electrons in Sodium = 11

Number of protons in sodium = 11

Number of neutrons in sodium = 12

Number of nucleons in sodium = 23

Solution 22:

Inert elements are the elements which have completely filled valence shell. Since, they are already stable and do not need more electrons, they do not combine with other atoms. So, they exist as monoatoms in molecule.

Solution 23:

1. **(a)** Hydrogen
2. **(b)** K shell.
3. **(c)** Magnesium
4. **(d)** Isotopes
5. **(e)** Helium has zero valency.

Solution 24:

1. **(a)** The number of protons = 9
2. **(b)** The number of neutrons = $19 - 9 = 10$
3. **(c)** The number of electrons = 9

Solution 25:

Atomic number is the number of protons of an atom which is unique to an atom but mass number is the total of number of protons and number of neutrons which may or may not be same to other atoms since there is probability of combination of number of protons and number of neutrons be same for two atoms.