CLASS –IX ASSINGMENT --- JULY2010
MULTIPLE CHOICE QUESTION ON PRACTICLE SKILLS
EXPERIMENT NO --18

1. Which one of the following does not form true solution in water
a. common salt b. alum c. albumin d. cane sugar

2. Which one of the following forms a colloid in water
a . common salt b. alum c. albumin d. cane sugar

3.. Which one of the following forms a colloid in water
a. soil b. sand c. chalk powder d .none of these

4. Which one of the following is highly soluble in water
a. common salt b. potash alum c. cane sugar d. starch

5. Which one of the following does not form true solution in water at room temperature
. common salt b. alum c. albumin d. cane sugar

6. . Which one of the following is least soluble in water
a. common salt b. potash alum c. cane sugar d. starch

7.Which one of the following can be separated by filtration
a. common salt and water b. sugar and water
c. soil and water d. starch and water

8. In which one of the following the particle size is in the range of 0.1 nm to 1.0 nm
\\\a. true solution b . suspension c. colloid d. none of these

9. .. Which one of the following forms a colloid in water
a. sugar b. sand c. chalk powder d. none of these

10.Tyndall effect is observed in which one of the following
a. true solution b . suspension c. colloid d. none of these

11.Sedimention does not take place in which one of the following
a. common salt and water b. sugar and water
c. soil and water d. starch and water

12.Which one will boil at a constant temperature ?
a. true solution b. colloid c. compound d. each one ..

13s.Which represents a constant composition
a. true solution b. colloid c. suspension d.. none of these

CLASS –IX ASSINGMENT --- JULY2010

MULTIPLE CHOICE QUESTION ON PRACTICLE SKILLS

EXPERIMENT NO –2

1. Which one of the following is an example of heterogeneous mixture
a. NaCl + water b. kerosene + water c. sugar + water d. alum + water

2. Which one of the following is a mixture of two elements
a. iron fillings + sulphur powder b. sugar + common salt
c. water +milk d. air

3. Which one is not the property of a mixture
a. heterogeneous b. variable composition c . constant composition d. none of these

4.Correct method of separating iron fillings from sulphur powder
a. heating in china dish b. heating in boiling tube c. heating in water d. moving magnet

5. In the formation of which one a large quantity of heat is involved
a. compound b. colloid c. true solution d. each one

6. When a magnet is moved through of iron fillings and sulphur powder , then
a. iron fillings will cling to magnet b. black mass is produced
c. sulphur will cling to magnet d. none of these

7.When a mixture of iron fillings is heated with sulphur powder
a. nothing happens b. black mass is formed c. yellow mass is formed d. white mass is formed

8. which one is a mixture
a. salad of fruits b. air c. sea water d. all of the above

9. Name the process of conversion of solid to vapor
a. evaporation b. fusion c. distillation d. sublimation

10.Which one of the following is a pure substance
a.milk b.cold drink c .carbon di oxide d. brass

11.What is observed when iron fillings and sulphur powder is heated in a china dish
a. sulphur melts b. iron melts c. mixture evaporates d.. yellow mass is formed

12. Mixture is always –
a. homogenous b. heterogeneous c. homogenous or heterogeneous d. none of these

13. Compound is always –
a. homogenous b. heterogeneous c. homogenous or heterogeneous d. none of these

14..Which of the following is a compound
a. blood b. air c. soil d. water

15.On heating iron and sulphur in a china dish
a. mixture is formed b. compound is formed c. both are formed d. no change

CLASS –IX ASSINGMENT --- AUG2010

MULTIPLE CHOICE QUESTION ON PRACTICLE SKILLS
EXPERIMENT NO –3

1.The color of copper sulphate solution is
a. green b. blue c. white d. brown

2. The reaction is a –
$BaCl_2 + Na_2SO_4$ ---------- $BaSO_4 + 2NaCl$
a.displacement reaction b. decomposition reaction c. double displacement reaction d. combination reaction

3 . I n the reaction , the colour of fumes of nitrogen di oxide is –
$2Pb(NO_3)_2$ ---------- $2PbO + 4NO_2 + O_2$

3. The reaction is a
$Fe + CuSO_4$ ---------- $FeSO_4 + Cu$
a.displacement reaction b. decomposition reaction c. double displacement reaction d. combination reaction

4.In which of the following reaction , a more active metal displaces less reactive metal
a.displacement reaction b. decomposition reaction c. double displacement reaction d. combination reaction

5.The reaction is a –

a.displacement reaction b. decomposition reaction c. double displacement reaction d. combination reaction

6.The gas evolved on adding dil. H_2SO_4 in some pieces of Zn is

a. oxygen b. hydrogen sulphide c. hydrogen d. sulphur di oxide

7.On mixing sodium sulphate and barium chloride
a. yellow ppt is formed b.white ppt is formed c. black ppt is formed d. grey ppt is formed

8. On burning magnesium ribbon
a. white substance is formed b .black substance is formed c. yellow substance is formed d. grey substance is formed

9. The reaction is a

$Zn + H_2SO_4$ ---------- $ZnSO_4 + H_2$

a.displacelacement reaction b. decomposition reaction c. double displacement reaction d. combination reaction

10.The color of residue left behind on heating lead nitrate (hot)

a. green b. white c. reddish brown d light yellow

11. The colour of nitrogen di oxide is

a. black b. orange c. blue d. reddish brown

12. The insoluble precipitate formed when barium chloride and sodium sulphate is mixed Is of –

a. sodium chloride b. barium sulphate c. barium sulphite d. none of these

13. The colour of insoluble precipitate formed when barium chloride and sodium sulphate is mixed Is

a. blue b. white c. black d. yellow

14. Burning of magnesium in oxygen is a

a.displacelacement reaction b. decomposition rection c. double displacement reaction d. combination reaction

15. The rectionwhich takes place when magnesium burns in air is

A $Mg + O_2$ -------- MgO

B. $Mg + N_2$ ------- Mg_3N_2

C. both a and b

d. none of these

CLASS –IX ASSINGMENT --- AUG2010

MULTIPLE CHOICE QUESTION ON PRACTICLE SKILLS
EXPERIMENT NO –4

1.Which one of the following can be separated by sublimation
a. camphor b. sugar c. salt d. copper sulphate

2 Which one of the following can be separated by sublimation
a. iodine b. sugar c. salt d. iron sulphate

4. Name the process of conversion of solid into vapor
a. sublimation b. evaporation c. condensation d. fusion

5. If sand is added to water –
a. it dissolves b. remains insoluble c. dissolves slightly d. none of these

6. The mixture of sand and ammonium chloride is separated by

a. sedimentation b. sublimation c. evaporation d. both a and b

7. Choose the homogenous mixture
a. sand and water b. sugar and water c. chalk and water d.. none

8. Salt can be separated from water by
a. evaporation b. distillation c.. condensation d. both a and b

9.The sand during filtration
a. remains on filter paper b. passes through filter paper c. both a and b d . none

10.Which of the following sublimes during heating
a. dry sait b. dry ice c ice d. all of them

CLASS –IX ASSINGMENT --- SEPSS2010

MULTIPLE CHOICE QUESTION ON PRACTICLE SKILLS
EXPERIMENT NO –5

1.What is the melting point of ice
a. 100K b. 273K c. 373K d.0 K

2.When does ice and water together exist at atmospheric pressure

a. at 0C b. below 0C c. above 0C d. never

3.Which one is true about kinetic energy of molecules of ice and water at 0C
a..it is high in ice b. it is high in water c. it is same in both d.. none

4.What is the boiling point of pure water
a. 273K b. 373 K c. 173 K d. 200K

5.Which of the following will increase the rate of evaporation of water
a. surface area b. temperature c.. wind energy d. all of the above

6.The heat energy needed to convert 1Kg of water into steam is called
a. latent heat of vaporization b. latent heat of fusion c. molar heat d.. specific heat

7.Heat is liberated when
a. water boils b. ice melts c. vapor condenses d. both a and c

8. During solidification process of water into ice
a. temperature is decreased b. temperature is increased c. heat is liberated d. heat is absorbed

9. When ice converts into water
a. temperature is decreased b. temperature is increased c. heat is liberated d. heat is absorbed

10.On supplying heat to a solid ,its temperature may
a. decrease b. increase c. remain constant d. increase then decrease

CLASS: IX
CHEMISTRY
ASSIGNMENT NO: 1

MATTER IN OUR SURROUNDING **MONTH: JULY – 2010**

1. What is matter? What are the physical states of matter?
2. With the help of an activity, show particulate nature of matter.
3. What are characteristics of particles of matter?
4. What is diffusion? Give an example.
5. Write any five characteristics of solids, liquids and gases.
6. Why can we smell hot food from a distance?
7. Why does a solid change into liquid on heating?
8. Define latent heat of fusion.
9. Give reason – A gas fills completely the vessel in which it is kept.
10. Why a wooden table should be called solid?
11. Why more serious burns are caused by steam at 100^0c than water at same temperature?
12. Why can a sponge be compressed though it is a solid?
13. Carry out following conversion (a) 50^0c to Kelvin (b) 200 k to Celsius.
14. Name the conditions to liquefy a gas.
15. Why are clothes spread out for drying?
16. What kind of clothes is most suitable for summers? Why?
17. What is sublimation? Explain with help of an activity and a diagram.
18. Washed clothes dry up more quickly on a hot summer day than on a rainy day. Why?
19. What is the difference between evaporation and boiling?
20. Define latent heat of vaporization.
21. Wearing synthetic clothes in summers is usually avoided. Why?
22. How does water get cooled in an earthen pot?
23. What is dry ice?
24. Name the factors affecting evaporation.
25. How does evaporation cause cooling? Explain with example.
26. How can physical state of matter be changed?

CLASS ---1X

CHEMISTRY – ASSIGNMENT 2 MONTH- AUG -SEP

– ITS MATTER AROUND US PURE

1. What are mixtures? Give examples.
2. What are pure substances?
3. Why do we call sugar a pure substance?
4. What are saturated and unsaturated solutions?
5. Define a solution.
6. What is a suspension? Give its example and properties.
7. Define concentration of a solution.
8. What is Tyndall effect?
9. What is the difference between True solution and colloids?
10. What are alloys? Why are alloys called as mixture?
11. Write the characteristics of brass.
12. Define solute and solvent.
13. What is solubility?
14. Give properties of a true solution.
15. Why do we need to separate mixtures?
16. How can we separate cream from milk?
17. Write the applications of centrifugation.
18. How can we separate a mixture of salt and ammonium chloride? Draw a diagram.
19. What is chromatography? Explain the process.
20. How can you separate copper sulphate from an impure sample?
21. What types of mixture are separated by fractional distillation process? Draw a diagram.
22. Draw a diagram to explain the process of separating acetone and water.
23. Explain with the diagram the method of separating kerosene oil from water by using a separating funnel.
24. What are physical and chemical changes?
25. Write difference between mixtures and compounds.
26. Give definitions of elements and compounds.
27. What are metals and non-metals?
28. If 110g of salt is present in 550 g of solution, calculate the concentration.
29. What is the concentration of a solution which contains 16 g of urea in 120 g of solutions?
30. How much water should be added to 15 g of salt to obtain 15% salt solution?
31. A solution contains 5.6 ml of alcohol mixed with 76 ml of water. Calculate the concentration of the solution.
32. How much water should be added to 12 ml of alcohol to obtain 12% alcohol solution?
33. If 25 ml of acetone is present in 150 ml of its aqueous solution, calculate the concentration.