CHAPTER 13

Hydrocarbons

13.2 Alkanes

1. Which of the following alkane cannot be made in good yield by Wurtz reaction?

(a) *n*-Hexane (b) 2, 3-Dimethylbutane
(c) *n*-Heptane (d) *n*-Butane

(NEET 2020)

2. The alkane that gives only one monochloro product on chlorination with Cl_2 in presence of diffused sunlight is

(a) 2,2-dimethylbutane (b) neopentane
(c) *n*-pentane (d) isopentane.

(Odisha NEET 2019)

3. Hydrocarbon (*A*) reacts with bromine by substitution to form an alkyl bromide which by Wurtz reaction is converted to gaseous hydrocarbon containing less than four carbon atoms. (*A*) is

(a) $CH \equiv CH$ (b) $CH_2 = CH_2$
(c) $CH_3 — CH_3$ (d) CH_4 *(NEET 2018)*

4. With respect to the conformers of ethane, which of the following statements is true?

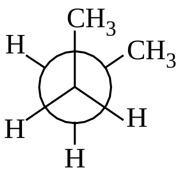
(a) Bond angle changes but bond length remains same.
(b) Both bond angle and bond length change.
(c) Both bond angle and bond length remain same.
(d) Bond angle remains same but bond length changes. *(NEET 2017)*

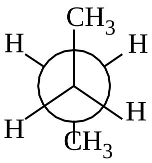
5. The correct statement regarding the comparison of staggered and eclipsed conformations of ethane, is

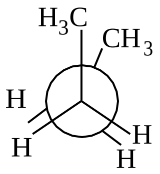
(a) the eclipsed conformation of ethane is more stable than staggered conformation even though the eclipsed conformation has torsional strain
(b) the staggered conformation of ethane is more stable than eclipsed conformation, because staggered conformation has no torsional strain
(c) the staggered conformation of ethane is less stable than eclipsed conformation, because staggered conformation has torsional strain
(d) the eclipsed conformation of ethane is more stable than staggered conformation, because eclipsed conformation has no torsional strain.

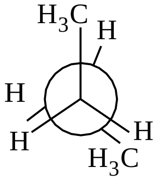
(NEET-I 2016)

6. In the following the most stable conformation of *n*-butane is

(a)


(b)


(c)


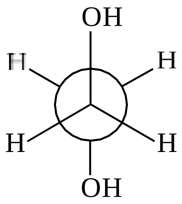
(d)


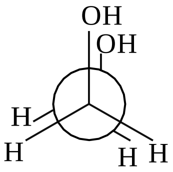
(2010)

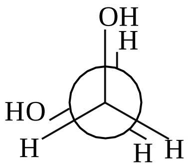
7. Liquid hydrocarbons can be converted to a mixture of gaseous hydrocarbons by

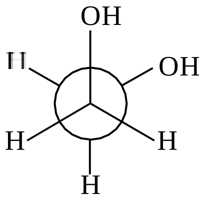
(a) oxidation
(b) cracking
(c) distillation under reduced pressure
(d) hydrolysis. *(2010)*

8. Which of the following conformers for ethylene glycol is most stable?

(a)


(b)


(c)


(d)


(Mains 2010)

9. Dihedral angle in staggered form of ethane is
(a) 0° (b) 120°
(c) 60° (d) 180° *(2000)*

10. Which of the following reaction is expected to readily give a hydrocarbon product in good yields?
(a) $CH_3CH_3 \xrightarrow[h\upsilon]{Cl_2}$ (b) $(CH_3)_2CHCl \xrightarrow{C_2H_5OH}$
(c) $RCOOK \xrightarrow[\text{Oxidation}]{\text{Electrolysis}}$ (d) $RCOOAg \xrightarrow{I_2}$
(1997)

11. In the commercial gasolines, the type of hydrocarbons which are more desirable is
(a) linear unsaturated hydrocarbon
(b) toluene
(c) branched hydrocarbon
(d) straight-chain hydrocarbon. *(1997)*

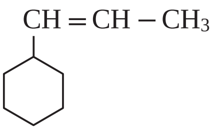
12. The most stable conformation of *n*-butane is
(a) gauche (b) staggered
(c) skew boat (d) eclipsed. *(1997)*

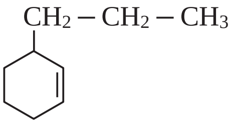
13. Which of the following is used as an antiknocking material?
(a) Glyoxal (b) Freon
(c) T.E.L. (d) Ethyl alcohol *(1996)*

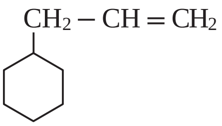
14. Reactivity of hydrogen atoms attached to different carbon atoms in alkanes has the order
(a) tertiary > primary > secondary
(b) primary > secondary > tertiary
(c) both (a) and (b)
(d) tertiary > secondary > primary. *(1993)*

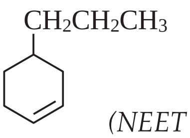
13.3 Alkenes

15. An alkene on ozonolysis gives methanal as one of the product. Its structure is

(a) 

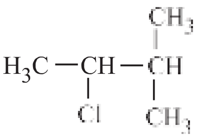
(b) 

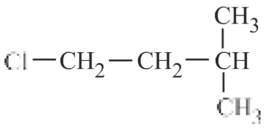
(c) 

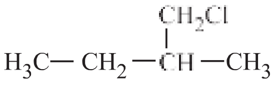
(d) 

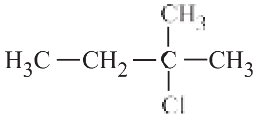
(NEET 2020)

16. An alkene *A* on reaction with O_3 and Zn–H_2O gives propanone and ethanal in equimolar ratio. Addition of HCl to alkene *A* gives *B* as the major product. The structure of product *B* is

(a) 
$H_3C-\underset{Cl}{CH}-\underset{CH_3}{\overset{CH_3}{CH}}$

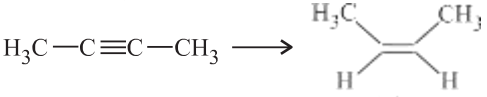
(b) 
$Cl-CH_2-CH_2-\underset{CH_3}{\overset{CH_3}{CH}}$

(c) 
$H_3C-CH_2-\overset{CH_2Cl}{CH}-CH_3$

(d) 
$H_3C-CH_2-\underset{Cl}{\overset{CH_3}{C}}-CH_3$

(NEET 2019)

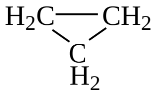
17. The most suitable reagent for the following conversion, is



cis-2-butene

(a) Hg^{2+}/H^+, H_2O (b) Na/liquid NH_3
(c) H_2, Pd/C, quinoline (d) Zn/HCl *(NEET 2019)*

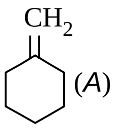
18. Which of the following compounds shall not produce propene by reaction with HBr followed by elimination or direct only elimination reaction?

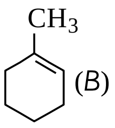
(a) 

(b) $H_3C-\overset{H_2}{C}-CH_2OH$
(c) $H_2C=C=O$
(d) $H_3C-\overset{H_2}{C}-CH_2Br$ *(NEET-II 2016)*

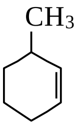
19. The compound that will react most readily with gaseous bromine has the formula
(a) C_3H_6 (b) C_2H_2
(c) C_4H_{10} (d) C_2H_4 *(NEET-II 2016)*

20. In the reaction with HCl, an alkene reacts in accordance with the Markovnikov's rule to give a product 1-chloro-1-methyl-cyclohexane. The possible alkene is

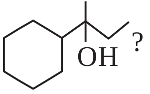
(a)  (*A*)

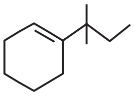
(b)  (*B*)

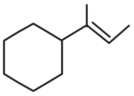
(c) (*A*) and (*B*)

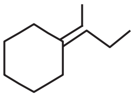
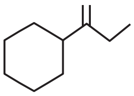
(d) 

(2015)

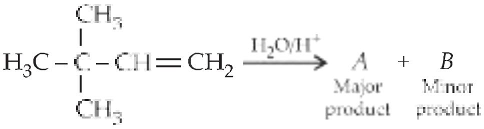
21. Which of the following is not the product of dehydration of  ?

(a) 

(b) 

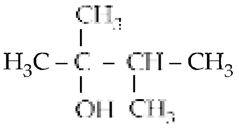
(c)  (d)  *(2015)*

22. In the following reaction

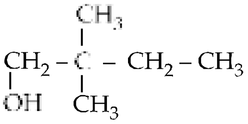


$$H_3C-\overset{CH_3}{\underset{CH_3}{C}}-CH=CH_2 \xrightarrow{H_2O/H^+} \underset{\text{Major product}}{A} + \underset{\text{Minor product}}{B}$$

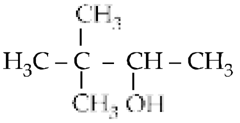
The major product is

(a) 

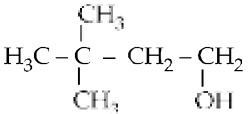
$H_3C-\overset{CH_3}{\underset{OH}{C}}-\underset{CH_3}{CH}-CH_3$

(b) 

$\underset{OH}{CH_2}-\overset{CH_3}{\underset{CH_3}{C}}-CH_2-CH_3$

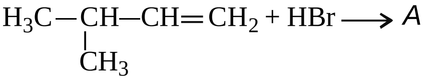
(c) 

$H_3C-\overset{CH_3}{\underset{CH_3}{C}}-\underset{OH}{CH}-CH_3$

(d)  *(2012)*

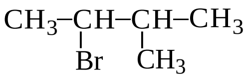
$H_3C-\overset{CH_3}{\underset{CH_3}{C}}-CH_2-\underset{OH}{CH_2}$

23. Which of the following compounds will exhibit *cis-trans* (geometrical) isomerism?
(a) Butanol (b) 2-Butyne
(c) 2-Butenol (d) 2-Butene *(2009)*

24. 

$H_3C-\underset{CH_3}{CH}-CH=CH_2 + HBr \longrightarrow A$

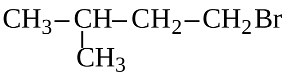
A (predominantly) is

(a) 

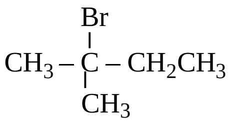
$CH_3-\underset{Br}{CH}-\underset{CH_3}{CH}-CH_3$

(b) 

$CH_3-\underset{CH_3}{CH}-\underset{Br}{CH}-CH_3$

(c) 

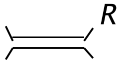
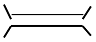
$CH_3-\underset{CH_3}{CH}-CH_2-CH_2Br$

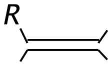
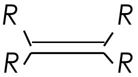
(d)  *(2008)*

$CH_3-\overset{Br}{\underset{CH_3}{C}}-CH_2CH_3$

25. Which of the compound with molecular formula C_5H_{10} yields acetone onozonolysis?
(a) 3-Methyl-1-butene (b) Cyclopentane
(c) 2-Methyl-1-butene (d) 2-Methyl-2-butene
(2007)

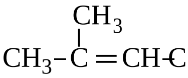
26. Which one of the following alkenes will react faster with H_2 under catalytic hydrogenation conditions?

(a)  (b) 

(c)  (d) 

(*R* = alkyl substituent)
(2005)

27. Reaction of HBr with propene in the presence of peroxide gives
(a) isopropyl bromide (b) 3-bromopropane
(c) allyl bromide (d) *n*-propyl bromide.
(2004)

28. The compound,  $CH_3-\overset{CH_3}{C}=CH-CH_3$ on reaction with $NaIO_4$ in the presence of $KMnO_4$ gives
(a) CH_3COCH_3
(b) $CH_3COCH_3 + CH_3COOH$
(c) $CH_3COCH_3 + CH_3CHO$
(d) $CH_3CHO + CO_2$ *(2003)*

29. Geometrical isomers differ in
(a) position of functional group
(b) position of atoms
(c) spatial arrangement of atoms
(d) length of carbon chain. *(2002)*

30. In preparation of alkene from alcohol using Al_2O_3 which is the effective factor?
(a) Porosity of Al_2O_3
(b) Temperature
(c) Concentration
(d) Surface area of Al_2O_3 *(2001)*

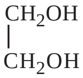
31. Which reagent converts propene to 1-propanol?
(a) H_2O, H_2SO_4
(b) B_2H_6, H_2O_2, OH^-
(c) $Hg(OAc)_2$, $NaBH_4/H_2O$
(d) Aq. KOH *(2000)*

32. Which is maximum stable?
(a) 1-Butene
(b) *cis*-2-Butene
(c) *trans*-2-Butene
(d) All have same stability. *(2000)*

33. 2-Butene shows geometrical isomerism due to
(a) restricted rotation about double bond
(b) free rotation about double bond
(c) free rotation about single bond
(d) chiral carbon. *(2000)*

34. 2-Bromopentane is heated with potassium ethoxide in ethanol. The major product obtained is
(a) *trans*-2-pentene (b) 1-pentene
(c) 2-ethoxy pentane (d) 2-*cis*-pentene.
(1998)

35. In a reaction, $CH_2=CH_2 \xrightarrow[\text{acid}]{\text{Hypochlorous}} M \xrightarrow{R}$



$\begin{array}{l} CH_2OH \\ | \\ CH_2OH \end{array}$

where, M = Molecule and R = Reagent. M and R are

(a) CH_3CH_2OH and HCl

(b) $CH_2=CH_2$ and heat

(c) CH_3CH_2Cl and NaOH

(d) CH_2Cl-CH_2OH and aq. $NaHCO_3$. *(1997)*

36. The reaction,

$CH_2=CH-CH_3 + HBr \longrightarrow CH_3CHBr-CH_3$ is

(a) electrophilic substitution

(b) free radical addition

(c) nucleophilic addition

(d) electrophilic addition. *(1996)*

37. Which of the following has zero dipole moment?

(a) 1-Butene (b) 2-Methyl-1-propene

(c) *cis*-2-Butene (d) *trans*-2-Butene *(1996)*

38. One of the following which does not observe the anti-Markownikoff's addition of HBr, is

(a) pent-2-ene (b) propene

(c) but-2-ene (d) but-1-ene. *(1994)*

39. Reduction of 2-butyne with sodium in liquid ammonia gives predominantly

(a) *cis*-2-butene (b) no reaction

(c) *trans*-2-butene (d) *n*-butane. *(1993)*

40. The restricted rotation about carbon carbon double bond in 2-butene is due to

(a) overlap of one *s* and sp^2-hybridized orbitals

(b) overlap of two sp^2-hybridized orbitals

(c) overlap of one *p* and one sp^2-hybridized orbitals

(d) sideways overlap of two *p*-orbitals. *(1993)*

41. Which one of the following can exhibit *cis-trans* isomerism?

(a) $CH_3-CHCl-COOH$

(b) $H-C\equiv C-Cl$

(c) $ClCH=CHCl$

(d) $ClCH_2-CH_2Cl$ *(1989)*

13.4 Alkynes

42. In the following reaction,

$H_3C-C\equiv CH \xrightarrow[\text{873 K}]{\text{red hot iron tube}} A,$

the number of sigma(σ) bonds present in the product A, is

(a) 21 (b) 9

(c) 24 (d) 18

(Odisha NEET 2019)

43. Which one is the correct order of acidity?

(a) $CH\equiv CH > CH_3-C\equiv CH > CH_2=CH_2 > CH_3-CH_3$

(b) $CH\equiv CH > CH_2=CH_2 > CH_3-C\equiv CH > CH_3-CH_3$

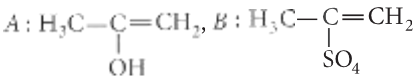
(c) $CH_3-CH_3 > CH_2=CH_2 > CH_3-C\equiv CH > CH\equiv CH$

(d) $CH_2=CH_2 > CH_3-CH=CH_2 > CH_3-C\equiv CH > CH\equiv CH$

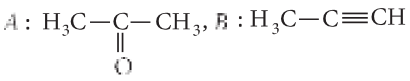
(NEET 2017)

44. Predict the correct intermediate and product in the following reaction :

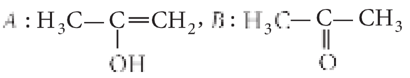
$H_3C-C\equiv CH \xrightarrow[HgSO_4]{H_2O,\ H_2SO_4} \underset{(A)}{\text{Intermediate}} \rightarrow \underset{(B)}{\text{Product}}$

(a) 

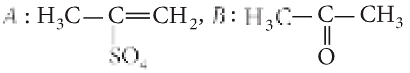
$A : H_3C-\underset{\underset{OH}{|}}{C}=CH_2,\ B : H_3C-\underset{\underset{SO_4}{|}}{C}=CH_2$

(b) 

$A : H_3C-\underset{\underset{O}{\|}}{C}-CH_3,\ B : H_3C-C\equiv CH$

(c) 

$A : H_3C-\underset{\underset{OH}{|}}{C}=CH_2,\ B : H_3C-\underset{\underset{O}{\|}}{C}-CH_3$

(d) 

$A : H_3C-\underset{\underset{SO_4}{|}}{C}=CH_2,\ B : H_3C-\underset{\underset{O}{\|}}{C}-CH_3$

(NEET 2017)

45. The pair of electrons in the given carbanion, $CH_3C\equiv C^-$, is present in which of the following orbitals?

(a) sp^2 (b) sp

(c) $2p$ (d) sp^3 *(NEET-I 2016)*

46. In the reaction

$H-C\equiv CH \xrightarrow[\text{(ii) } CH_3CH_2Br]{\text{(i) } NaNH_2/\text{liq.} NH_3} X \xrightarrow[\text{(ii) } CH_3CH_2Br]{\text{(i) } NaNH_2/\text{liq. } NH_3} Y,$

X and Y are

(a) X = 2-butyne, Y = 2-hexyne

(b) X = 1-butyne, Y = 2-hexyne

(c) X = 1-butyne, Y = 3-hexyne

(d) X = 2-butyne, Y = 3-hexyne. *(NEET-I 2016)*

47. Which of the following organic compounds has same hybridization as its combustion product (CO_2)?

(a) Ethane (b) Ethyne

(c) Ethene (d) Ethanol *(2014)*

48. In the following reaction :

$HC\equiv CH \xrightarrow[Hg^{2+}]{H_2SO_4} \text{'}P\text{'}$

Product '*P*' will not give

(a) Tollens' reagent test

(b) Brady's reagent test

(c) Victor Meyer test

(d) Iodoform test. *(Karnataka NEET 2013)*

49. Which of the following reagents will be able to distinguish between 1-butyne and 2-butyne?

(a) $NaNH_2$ (b) HCl

(c) O_2 (d) Br_2 *(Mains 2012)*

50. Considering the state of hybridization of carbon atoms, find out the molecule among the following which is linear?

(a) $CH_3 - CH = CH - CH_3$

(b) $CH_3 - C \equiv C - CH_3$

(c) $CH_2 = CH - CH_2 - C \equiv CH$

(d) $CH_3 - CH_2 - CH_2 - CH_3$ *(2011)*

51. Base strength of

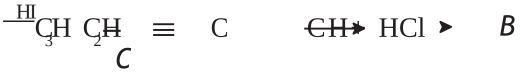
$\underset{(i)}{H_3C\bar{C}H_2}$, $\underset{(ii)}{H_2C = \bar{C}H}$ and $\underset{(iii)}{H - C \equiv \bar{C}}$

is in the order of

(a) (i) > (iii) > (ii) (b) (i) > (ii) > (iii)

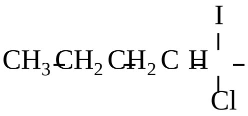
(c) (ii) > (i) > (iii) (d) (iii) > (ii) > (i) *(2008)*

52. Predict the product *C* obtained in the following reaction of 1-butyne.

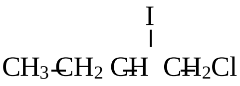


$\overline{\text{HI}}\,CH_3CH_2 \equiv C \quad CH \rightarrow HCl \rightarrow B$

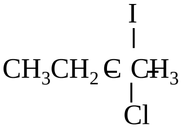
C

(a) 

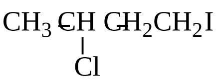
$CH_3CH_2CH_2\overset{I}{\underset{Cl}{C}}H$

(b) 

$CH_3CH_2\overset{I}{C}HCH_2Cl$

(c) 

$CH_3CH_2\overset{I}{\underset{Cl}{C}}CH_3$

(d)  *(2007)*

$CH_3\underset{Cl}{C}HCH_2CH_2I$

53. Products of the following reaction:

$CH_3C \equiv CCH_2CH_3 \xrightarrow[\text{(ii) hydrolysis}]{\text{(i) } O_3}$

are

(a) $CH_3COOH + CO_2$

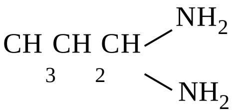
(b) $CH_3COOH + HOOCCH_2CH_3$

(c) $CH_3CHO + CH_3CH_2CHO$

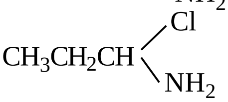
(d) $CH_3COOH + CH_3COCH_3$ *(2005)*

54. When $CH_3CH_2CHCl_2$ is treated with $NaNH_2$, the product formed is

(a) $CH_3 - CH = CH_2$ (b) $CH_3 - C \equiv CH$

(c) 

$CH_3CH_2CH(NH_2)_2$

(d)  *(2002)*

$CH_3CH_2CH(Cl)(NH_2)$

55. When acetylene is passed through dil. H_2SO_4 in the presence of $HgSO_4$, the compound formed is

(a) acetic acid (b) ketone

(c) ether (d) acetaldehyde. *(1999)*

56. The cylindrical shape of an alkyne is due to

(a) two sigma C – C and one π C – C bonds

(b) one sigma C – C and two π C – C bonds

(c) three sigma C – C bonds

(d) three π C – C bonds. *(1997)*

57. $R - CH_2 - CCl_2 - R \xrightarrow{\text{Reagent}} R - C \equiv C - R$

The reagent is

(a) Na (b) HCl in H_2O

(c) KOH in C_2H_5OH (d) Zn in alcohol. *(1993)*

58. A compound is treated with $NaNH_2$ to give sodium salt. Identify the compound.

(a) C_2H_2 (b) C_6H_6

(c) C_2H_6 (d) C_2H_4 *(1993)*

59. The shortest C–C bond distance is found in

(a) diamond (b) ethane

(c) benzene (d) acetylene. *(1991)*

60. Acetylenic hydrogens are acidic because

(a) sigma electron density of C – H bond in acetylene is nearer to carbon, which has 50% *s*-character

(b) acetylene has only open hydrogen in each carbon

(c) acetylene contains least number of hydrogens among the possible hydrocarbons having two carbons

(d) acetylene belongs to the class of alkynes with molecular formula, C_nH_{2n-2}. *(1989)*

61. Which is the most suitable reagent among the following to distinguish compound (3) from rest of the compounds?

(1) $CH_3 - C \equiv C - CH_3$

(2) $CH_3 - CH_2 - CH_2 - CH_3$

(3) $CH_3 - CH_2C \equiv CH$

(4) $CH_3 - CH = CH_2$

(a) Bromine in carbon tetrachloride

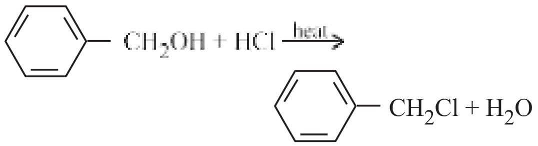
(b) Bromine in acetic acid

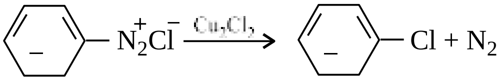
(c) Alk. $KMnO_4$

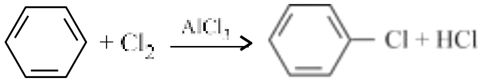
(d) Ammoniacal silver nitrate *(1989)*

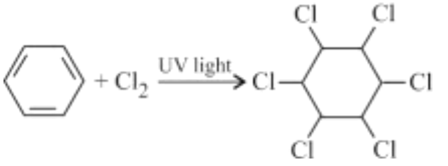
13.5 Aromatic hydrocarbons

62. Among the following the reaction that proceeds through an electrophilic substitution is

(a) 

(b) 

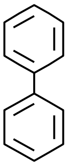
(c) 

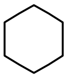
(d) 

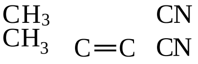
(NEET 2019)

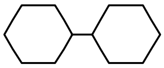
63. Which of the following can be used as the halide component for Friedel–Crafts reaction?
(a) Chlorobenzene (b) Bromobenzene
(c) Chloroethene (d) Isopropyl chloride
(NEET-II 2016)

64. In which of the following molecules, all atoms are coplanar?

(a) 

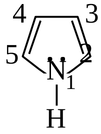
(b) 

(c) 
$\overset{CH_3}{\underset{CH_3}{}}C{=}C\overset{CN}{\underset{CN}{}}$

(d) 

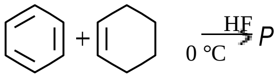
(NEET-II 2016)

65. In pyrrole the electron density is maximum on

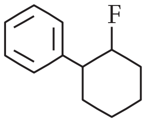


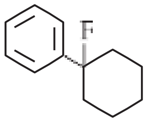
(a) 2 and 3 (b) 3 and 4
(c) 2 and 4 (d) 2 and 5
(NEET-II 2016)

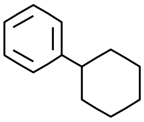
66. In the given reaction,

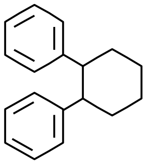


the product *P* is

(a) 

(b) 

(c) 

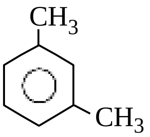
(d) 

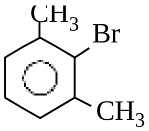
(NEET-II 2016)

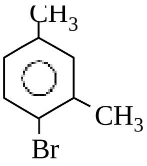
67. Consider the nitration of benzene using mixed conc. H_2SO_4 and HNO_3. If a large amount of $KHSO_4$ is added to the mixture, the rate of nitration will be
(a) unchanged (b) doubled
(c) faster (d) slower.
(NEET-I 2016)

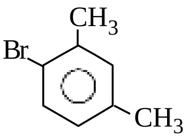
68. The oxidation of benzene by V_2O_5 in the presence of air produces
(a) maleic anhydride
(b) benzoic acid
(c) benzaldehyde
(d) benzoic anhydride. *(2015)*

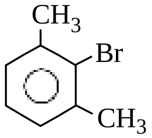
69. What products are formed when the following compound is treated with Br_2 in the presence of $FeBr_3$?

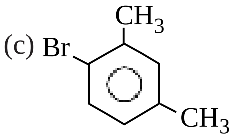


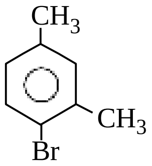
(a) 

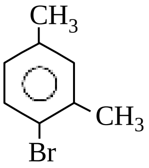
and


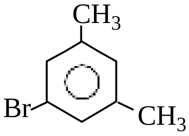
(b) 

and


(c) 

and


(d) 

and


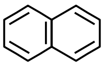
(2014)

70. Some *meta*-directing substituents in aromatic substitution are given. Which one is most deactivating?
(a) —COOH (b) $—NO_2$
(c) —C≡N (d) $—SO_3H$
(NEET 2013)

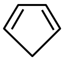
71. Which of the following compounds will not undergo Friedel-Crafts reaction easily?
(a) Nitrobenzene (b) Toluene
(c) Cumene (d) Xylene *(NEET 2013)*

72. Which of the following chemical system is non aromatic?

(a) 

(b) 

(c) 

(d) 

(Karnataka NEET 2013)

73. Among the following compounds the one that is most reactive towards electrophilic nitration is

(a) benzoic acid (b) nitrobenzene

(c) toluene (d) benzene.

(2012, 1992)

74. The reaction of toluene with Cl_2 in presence of $FeCl_3$ gives *X* and reaction in presence of light gives *Y*. Thus, *X* and *Y* are

(a) *X* = benzal chloride,
Y = *o*-chlorotoluene

(b) *X* = *m*-chlorotoluene,
Y = *p*-chlorotoluene

(c) *X* = *o*- and *p*-chlorotoluene,
Y = trichloromethyl benzene

(d) *X* = benzyl chloride,
Y = *m*-chlorotoluene.

(2010)

75. Benzene reacts with CH_3Cl in the presence of anhydrous $AlCl_3$ to form

(a) chlorobenzene (b) benzyl chloride

(c) xylene (d) toluene.

(2009)

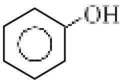
76. Nitrobenzene can be prepared from benzene by using a mixture of conc. HNO_3 and conc. H_2SO_4. In the mixture, nitric acid acts as a/an

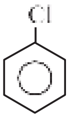
(a) acid (b) base

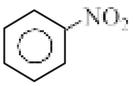
(c) catalyst (d) reducing agent.

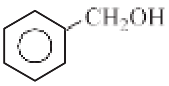
(2009)

77. Which one of the following is most reactive towards electrophilic attack?

(a) 

(b) 

(c) 

(d) 

(2008)

78. The order of decreasing reactivity towards an electrophilic reagent, for the following would be

(i) benzene (ii) toluene

(iii) chlorobenzene (iv) phenol

(a) (ii) > (iv) > (i) > (iii)

(b) (iv) > (iii) > (ii) > (i)

(c) (iv) > (ii) > (i) > (iii)

(d) (i) > (ii) > (iii) > (iv)

(2007)

79. Using anhydrous $AlCl_3$ as catalyst, which one of the following reactions produces ethylbenzene (PhEt)?

(a) $H_3C - CH_2OH + C_6H_6$

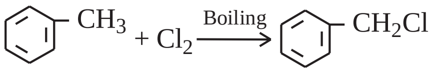
(b) $CH_3 - CH{=}CH_2 + C_6H_6$

(c) $H_2C{=}CH_2 + C_6H_6$

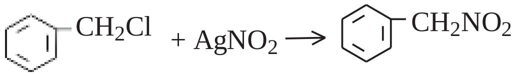
(d) $H_3C - CH_3 + C_6H_6$

(2004)

80. Which one of the following is a free-radical substitution reaction?

(a) 

(b) 

(c) 

(d) $CH_3CHO + HCN \rightarrow CH_3CH(OH)CN$ *(2003)*

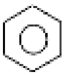
81. The correct order of reactivity towards the electrophilic substitution of the compounds aniline (I), benzene (II) and nitrobenzene (III) is

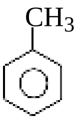
(a) III > II > I (b) II > III > I

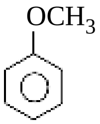
(c) I < II > III (d) I > II > III

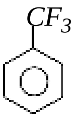
(2003)

82. Increasing order of electrophilic substitution for following compounds

(I) 

(II) 

(III) 

(IV) 

(a) IV < I < II < III (b) III < II < I < IV

(c) I < IV < III < II (d) II < III < I < IV

(2000)

83. In Friedel-Crafts reaction, toluene can be prepared by

(a) $C_6H_6 + CH_3Cl$ (b) $C_6H_5Cl + CH_4$

(c) $C_6H_6 + CH_2Cl_2$

(d) $C_6H_6 + CH_3COCl$

(2000)

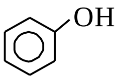
84. In Friedel-Crafts alkylation, besides $AlCl_3$ the other reactants are

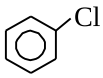
(a) $C_6H_6 + CH_3Cl$ (b) $C_6H_6 + CH_4$

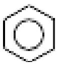
(c) $C_6H_6 + NH_3$ (d) $C_6H_6 + CH_3COCl$

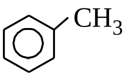
(1999)

85. Which of the following compounds will be most easily attacked by an electrophile?

(a) 

(b) 

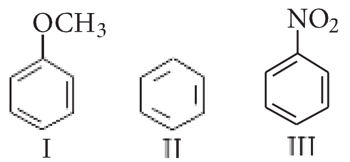
(c) 

(d) 

(1999, 1998)

86. Which one of these is not compatible with arenes?
(a) Electrophilic additions
(b) Delocalisation of π-electrons
(c) Greater stability
(d) Resonance *(1998)*

87. Among the following compounds (I-III) the correct order of reaction with electrophile is



(a) I > II > III (b) I = II > III
(c) II > III > I (d) III < I < II *(1997)*

88. Electrophile in the case of chlorination of benzene in the presence of $FeCl_3$ is
(a) Cl (b) $FeCl_3$
(c) Cl^+ (d) Cl^- *(1996)*

89. The reactive species in the nitration of benzene is
(a) NO_3 (b) HNO_3
(c) NO_2^+ (d) NO_2^- *(1994)*

90. Which is the correct symbol relating the two Kekule structures of benzene?
(a) $\rightleftharpoons$ (b) $\longrightarrow$
(c) $=$ (d) $\longleftrightarrow$ *(1993)*

91. Select the true statement about benzene amongst the following
(a) because of unsaturation benzene easily undergoes addition
(b) there are two types of C – C bonds in benzene molecule
(c) there is cyclic delocalisation of π-electrons in benzene
(d) monosubstitution of benzene gives three isomeric products. *(1992)*

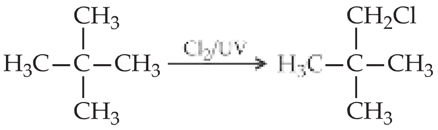
ANSWER KEY

1.	(c)	**2.**	(b)	**3.**	(d)	**4.**	(c)	**5.**	(b)	**6.**	(b)	**7.**	(b)	**8.**	(d)	**9.**	(c)	**10.**	(c)
11.	(c)	**12.**	(b)	**13.**	(c)	**14.**	(d)	**15.**	(c)	**16.**	(d)	**17.**	(c)	**18.**	(c)	**19.**	(a)	**20.**	(c)
21.	(a)	**22.**	(a)	**23.**	(d)	**24.**	(d)	**25.**	(d)	**26.**	(a)	**27.**	(d)	**28.**	(b)	**29.**	(c)	**30.**	(b)
31.	(b)	**32.**	(c)	**33.**	(a)	**34.**	(a)	**35.**	(d)	**36.**	(d)	**37.**	(d)	**38.**	(c)	**39.**	(c)	**40.**	(d)
41.	(c)	**42.**	(a)	**43.**	(a)	**44.**	(c)	**45.**	(b)	**46.**	(c)	**47.**	(b)	**48.**	(c)	**49.**	(a)	**50.**	(b)
51.	(b)	**52.**	(c)	**53.**	(b)	**54.**	(b)	**55.**	(d)	**56.**	(b)	**57.**	(c)	**58.**	(a)	**59.**	(d)	**60.**	(a)
61.	(d)	**62.**	(c)	**63.**	(d)	**64.**	(a)	**65.**	(d)	**66.**	(c)	**67.**	(d)	**68.**	(a)	**69.**	(c)	**70.**	(b)
71.	(a)	**72.**	(d)	**73.**	(c)	**74.**	(c)	**75.**	(d)	**76.**	(b)	**77.**	(a)	**78.**	(c)	**79.**	(c)	**80.**	(a)
81.	(d)	**82.**	(a)	**83.**	(a)	**84.**	(a)	**85.**	(a)	**86.**	(a)	**87.**	(a)	**88.**	(c)	**89.**	(c)	**90.**	(d)
91.	(c)																		

Hints & Explanations

1. (c) : Wurtz reaction is used for the preparation of higher alkanes containing even number of C-atoms. Thus this reaction cannot be used for the preparation of *n*-heptane.

2. (b) : In chlorination of alkanes, hydrogen is replaced by chlorine. In *neo*-pentane, only one type of hydrogen is present, hence replacement of any hydrogen atom will give the same product.



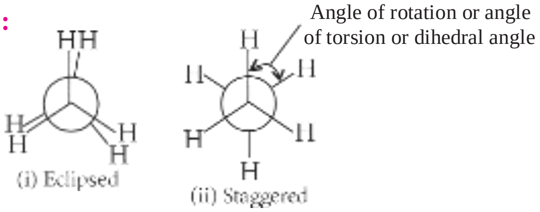
$$H_3C-\overset{\overset{\large CH_3}{|}}{\underset{\underset{\large CH_3}{|}}{C}}-CH_3 \xrightarrow{Cl_2/UV} H_3C-\overset{\overset{\large CH_2Cl}{|}}{\underset{\underset{\large CH_3}{|}}{C}}-CH_3$$

neo-pentane

3. (d) : $\underset{(A)}{CH_4} \xrightarrow{Br_2/h\nu} CH_3Br \xrightarrow{\text{Na/dry ether}}_{\text{}}$ $\overset{\text{Na/dry ether}}{\overset{\text{Wurtz reaction}}{\longrightarrow}} CH_3-CH_3$
(Less than 4 caron atoms)

4. (c) : Conformers of ethane have different dihedral angles but same bond angle and bond lengths.

5. (b) :



(i) Eclipsed (ii) Staggered

Newman's projections of ethane

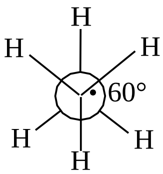
Magnitude of torsional strain depends upon the angle of rotation about C—C bond. Staggered form has the least torsional strain and the eclipsed form has the maximum torsional strain. So, the staggered conformation of ethane is more stable than the eclipsed conformation.

6. (b) : The anti-conformation is the most stable conformation of *n*-butane as in this, the bulky methyl groups are as far apart as possible thereby keeping steric repulsion at a minimum.

7. (b) : The process of cracking converts higher alkanes into smaller alkanes and alkenes. This process can be used for production of natural gas.

8. (d) : The conformation (d) is most stable because of intermolecular H-bonding.

9. (c) : The staggered form of ethane has the following structure and the dihedral angle is 60°, which means 'H' atoms are at an angle of 60° to each other.



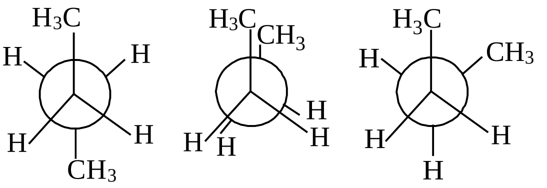
10. (c) : When an aqueous solution of sodium or potassium salt of carboxylic acid is electrolysed, hydrocarbon is evolved at anode.

$$2RCOOK \xrightarrow[\text{Oxidation}]{\text{Electrolysis}} \underset{\text{Anode}}{2RCOO^-} + \underset{\text{Cathode}}{2K^+}$$

At anode : $2RCOO^- - 2e^- \rightarrow \underset{\text{Alkane}}{R-R} + 2CO_2$

11. (c) : The branching of chain increases the octane number of a fuel. High octane number means better fuel.

12. (b) : Newman projections for *n*-butane are



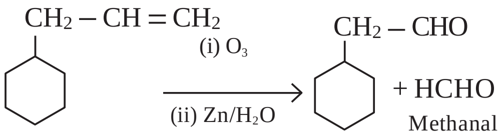
staggered eclipsed skew or gauche

The staggered conformation has minimum repulsion between the hydrogen atoms attached tetrahedrally to the two carbon atoms. Thus, it is the most stable conformation.

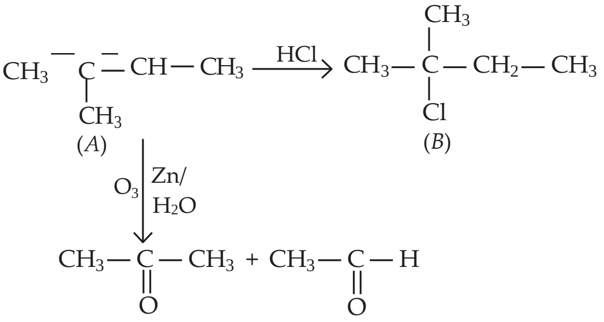
13. (c) : Tetraethyl lead $(C_2H_5)_4Pb$, is used as an antiknocking agent in gasoline used for running automobiles.

14. (d) : The reactivity of H-atom depends upon the stability of free radicals, therefore reactivity of H-atom follows the order : 3° > 2° > 1°.

15. (c) :

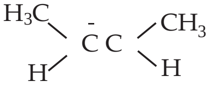


16. (d) :



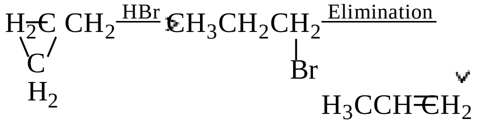
Addition of HCl to an alkene (*A*) will take place according to Markownikoff's rule.

17. (c) : $CH_3C \equiv CCH_3 \xrightarrow[\text{quinoline}]{H_2,\ Pd/C}$

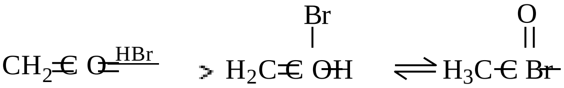


cis-2-Butene

18. (c) :



$CH_3CH_2CH_2OH \xrightarrow{HBr} \xrightarrow{\text{Elimination}} H_3C–CH = CH_2$

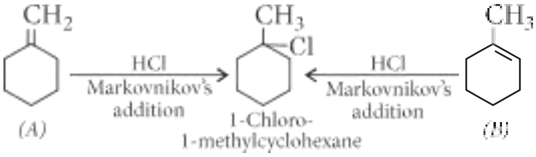


$$CH_2=C=O \xrightarrow{HBr} H_2C=\overset{Br}{\overset{|}{C}}-OH \rightleftharpoons H_3C-\overset{O}{\overset{||}{C}}-Br$$

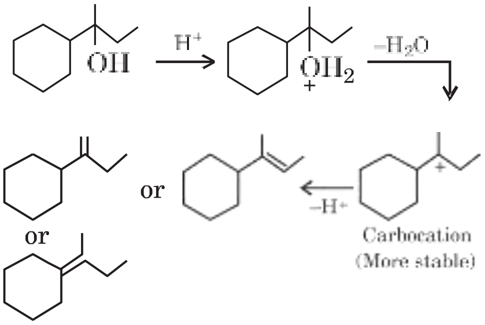
$CH_3CH_2CH_2Br \xrightarrow{\text{Elimination}} CH_3CH = CH_2$

19. (a) : Rate of free radical substitution with $Br_{2(g)}$ depends upon the stability of free radical. Propenyl free radical is allylic free radical which is more stable.

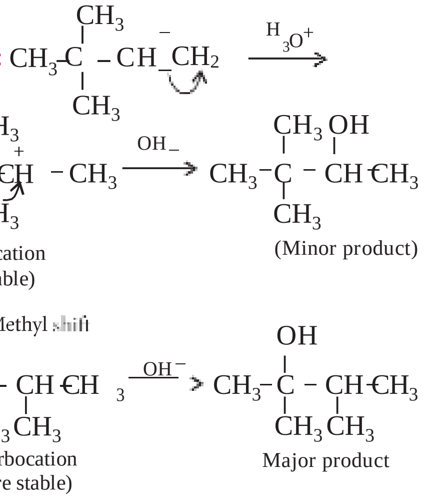
20. (c) :



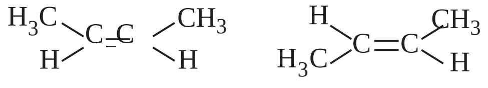
21. (a) :



22. (a) :

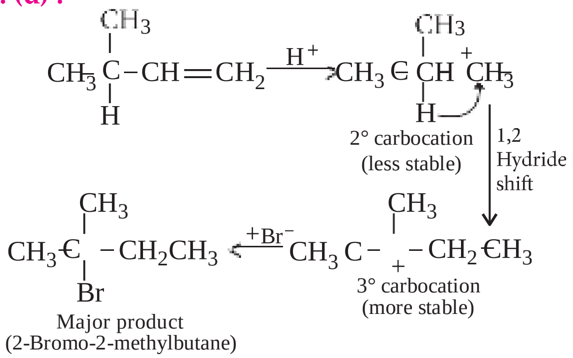


23. (d) : *Cis-trans* isomerism is exhibited by compounds having C = C, C = N and N = N groups, due to restricted rotation around the double bond. Among the given options, only 2-butene exhibits geometrical isomerism.

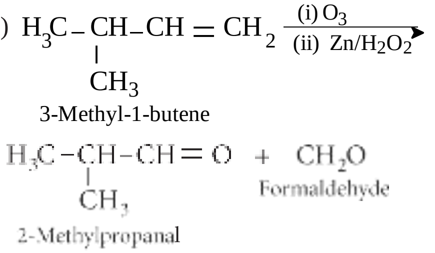


cis-form *trans*-form

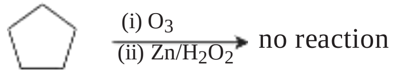
24. (d) :



25. (d) : (a)

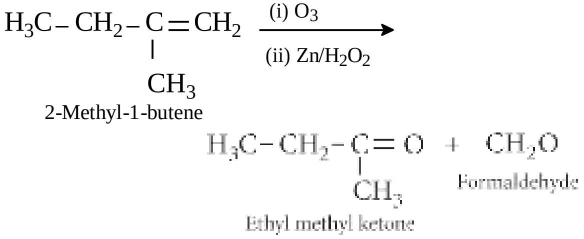


(b)

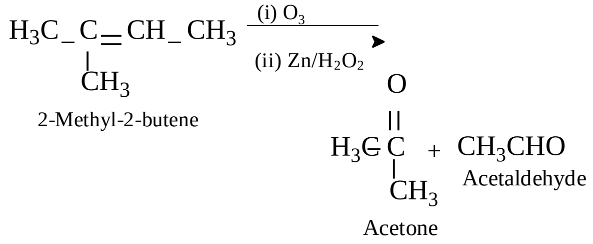


Cyclopentane

(c)

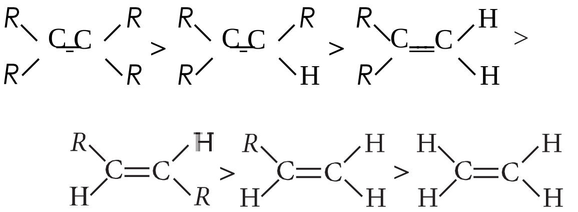


(d)

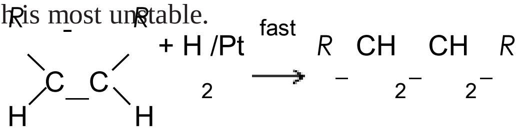


26. (a) : The relative rates of hydrogenation decrease with the increase of steric hindrance. Most stable the alkene, slowly it undergoes hydrogenation to give the product. Least substituted alkene is less stable and more reactive.

Order of stability is :



Hence, alkene which will react faster with H_2 is that which is most unstable.



27. (d) : The formation of *n*-propyl bromide in presence of peroxide can be explained as follow :

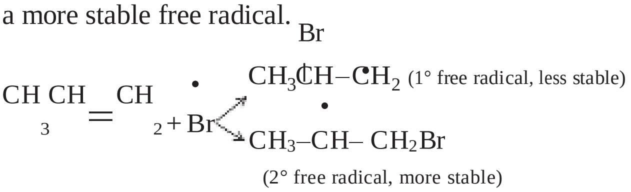
Step 1 : Peroxide undergoes fission to give free radicals.

$$R-O-O-R \longrightarrow 2R-\dot{O}$$

Step 2 : HBr combines with free radical to form bromine free radical.

$$R-\dot{O} + HBr \longrightarrow R-OH + \dot{B}r$$

Step 3 : $\dot{B}r$ attacks the double bond of the alkene to form a more stable free radical.



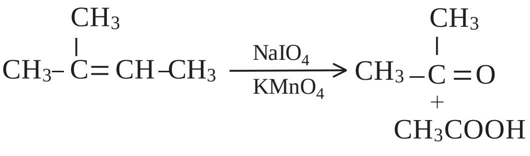
Step 4 : More stable free radical attacks on HBr.

$$CH_3\dot{C}HCH_2Br + HBr \longrightarrow CH_3CH_2CH_2Br + \dot{B}r$$

n-Propyl bromide

Step 5 : $\dot{B}r + \dot{B}r \longrightarrow Br_2$

28. (b) :



$$CH_3-\underset{}{\overset{CH_3}{C}}=CH-CH_3 \xrightarrow[KMnO_4]{NaIO_4} CH_3-\overset{CH_3}{C}=O + CH_3COOH$$

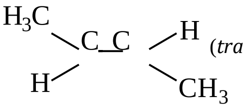
29. (c) : Geometrical isomers are those isomers which possess the same molecular and structural formula but differ in the arrangement of atoms or groups in space due to hindered rotation around the double bonded atoms.

30. (b) : Temperature is an effective factor because at different temperature it forms different products.

$$CH_3CH_2OH \xrightarrow[-H_2O]{Al_2O_3,\ 620\ K} CH_2=CH_2$$

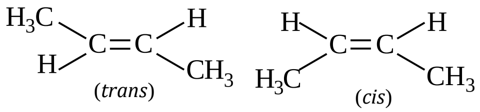
This is intramolecular dehydration. At lower temperature, intermolecular dehydration takes place between two molecules of alcohol and ether will be formed.

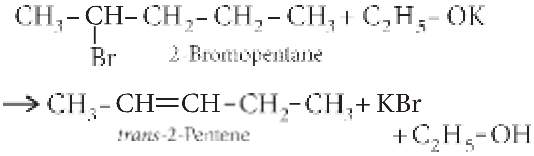
31. (b) : Propene adds to diborane (B_2H_6) giving an addition product. The addition compound on oxidation gives 1-propanol. Here addition of water takes place according to anti-Markownikoff's rule.

32. (c) :  (*trans*-2-butene)

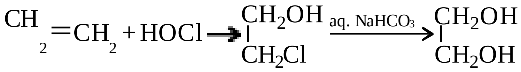
This is most stable as the repulsion between two methyl groups is least.

33. (a) : Due to restricted rotation about double bond, 2-butene shows geometrical isomerism.


(*trans*) (*cis*)

34. (a) : 

$$\underset{\text{2-Bromopentane}}{CH_3-\underset{\substack{|\\Br}}{CH}-CH_2-CH_2-CH_3} + C_2H_5-OK \longrightarrow \underset{\textit{trans}\text{-2-Pentene}}{CH_3-CH=CH-CH_2-CH_3} + KBr + C_2H_5-OH$$

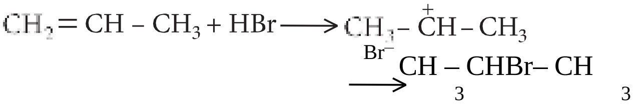
35. (d) : 

$$CH_2=CH_2 + HOCl \longrightarrow \begin{array}{l}CH_2OH\\|\\CH_2Cl\end{array} \xrightarrow{\text{aq. }NaHCO_3} \begin{array}{l}CH_2OH\\|\\CH_2OH\end{array}$$

Therefore, $M = CH_2Cl—CH_2OH$ and

R = aq. $NaHCO_3$

36. (d) : In this reaction, HBr undergoes heterolytic fission as $HBr \longrightarrow H^+ + Br^-$



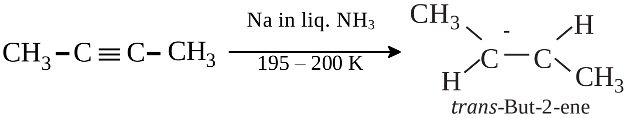
$$CH_2=CH-CH_3 + HBr \longrightarrow CH_3-\overset{+}{C}H-CH_3 \xrightarrow{Br^-} CH_3-CHBr-CH_3$$

This is an example of electrophilic addition reaction.

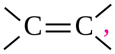
37. (d)

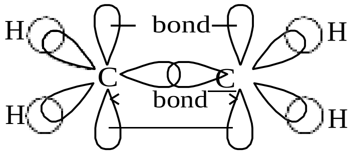
38. (c) : In case of but-2-ene ($CH_3-CH=CH-CH_3$) both double bonded carbons are identical. Therefore, it does not observe the anti-Markownikoff's addition of HBr.

39. (c) : Reduction of non-terminal alkynes with Na in liq. NH_3 at 195 - 200 K gives *trans*-2-butene.

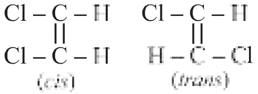


trans-But-2-ene
(Birch reduction)

40. (d) : 

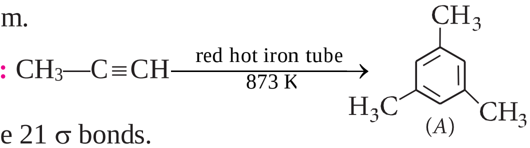


Restricted rotation is due to sideways overlap of two *p*-orbitals.

41. (c) : 

$$\underset{(\textit{cis})}{\begin{array}{c}Cl-C-H\\ \|\\ Cl-C-H\end{array}} \qquad \underset{(\textit{trans})}{\begin{array}{c}Cl-C-H\\ \|\\ H-C-Cl\end{array}}$$

1, 2-Dichloroethene exhibits *cis-trans* (geometrical) isomerism.

42. (a) : 

There are 21 σ bonds.

43. (a) : Alkanes, alkenes and alkynes follow the following trend in their acidic behaviour :

$$H\overset{sp}{C}\equiv\overset{sp}{C}H > H_2\overset{sp^2}{C}=\overset{sp^2}{C}H_2 > \overset{sp^3}{C}H_3—\overset{sp^3}{C}H_3$$

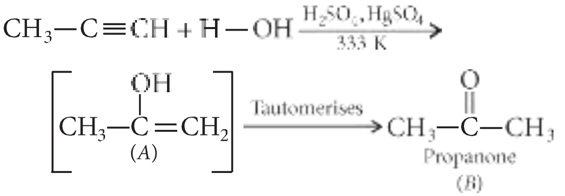
This is because *sp*-hybridised carbon is more electronegative than sp^2-hybridised carbon which is further more electronegative than sp^3-hybridised carbon. Hence, in ethyne proton can be released more easily than ethene and ethane.

Among alkynes the order of acidity is :

$$HC\equiv CH > CH_3—C\equiv CH > CH_3—C\equiv C—CH_3$$

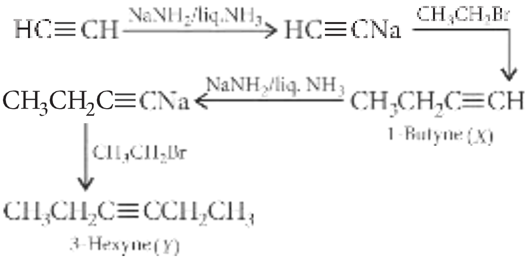
This is due to $+I$ effect of $—CH_3$ group.

44. (c) : In case of unsymmetrical alkynes addition of H_2O occurs in accordance with Markownikoff's rule.



45. (b) : $CH_3—\overset{sp}{C}\equiv\overset{sp}{C}^-$

Thus, pair of electrons is present in *sp*-hybridised orbital.

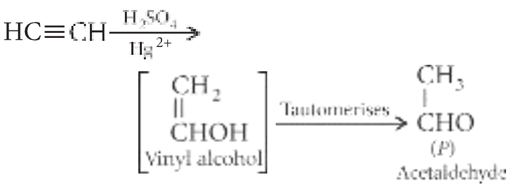
46. (c) : 

47. (b) : $C_2H_2 + \frac{5}{2}O_2 \longrightarrow 2CO_2 + H_2O$

Both ethyne and CO_2 have *sp*-hybridisation.

$O{=}\overset{sp}{C}{=}O \qquad H\overset{sp}{C}{\equiv}\overset{sp}{C}H$

48. (c) :



Acetaldehyde does not give Victor Meyer test.

49. (a) : Terminal alkynes (1-butyne) react with $NaNH_2$ to form sodium acetylide and evolve hydrogen but 2-butyne do not.

50. (b) : $CH_3-C\equiv C-CH_3$

In case of sp^3 hybridised carbon, bond angle is 109° 28′; sp^2 hybridised carbon, bond angle is 120° and *sp* hybridised carbon, bond angle is 180°.

So, only $CH_3-\underset{sp}{\underset{\downarrow}{C}}\equiv\underset{sp}{\underset{\downarrow}{C}}-CH_3$ is linear.

51. (b) : $\underset{sp}{H-C}\equiv\underset{sp}{C}-H > \underset{sp^2}{CH_2}=\underset{sp^2}{CH_2} > \underset{sp^3}{CH_3}\underset{sp^3}{CH_3}$

(Acidic character)

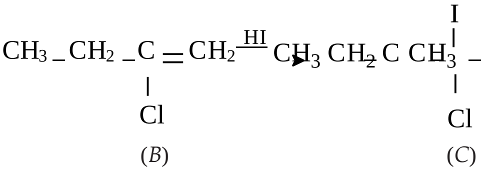
Conjugate base of the given acid :

$\bar{C}\equiv C-H < \bar{C}H=CH_2 < \bar{C}H_2CH_3$

(Basic character)

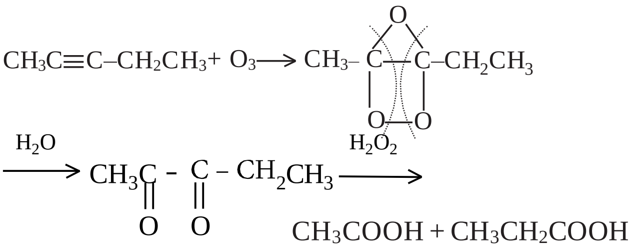
Conjugate base of stronger acid is weaker and vice-versa.

52. (c) : $CH_3-CH_2-C\equiv CH + HCl \longrightarrow$



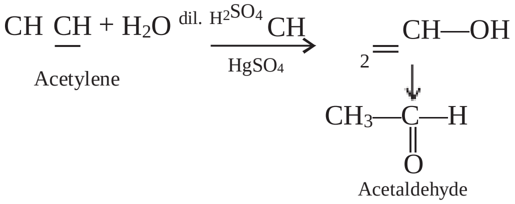
According to Markownikoff's rule, during hydrohalogenation to unsymmetrical alkene, the negative part of the addendum adds to less hydrogenated (*i.e.* more substituted) carbon atom.

53. (b) : On ozonolysis, higher alkynes form diketones which are further oxidised to dicarboxylic acid.



54. (b) : $CH_3CH_2CHCl_2 \xrightarrow{NaNH_2} CH_3C\equiv CH$

55. (d) :



56. (b) : In alkyne, two carbon atoms constituting the triple bond are *sp*-hybridised. Carbon undergoes *sp*-hybridisation to form two *sp*-hybrid orbitals. The two 2*p*-orbitals remain unhybridised. Hybrid orbitals form one sigma bond while two π-bonds are formed by unhybridised orbitals.

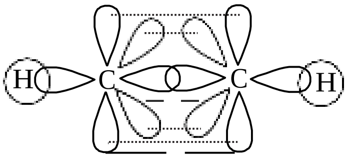
57. (c) : A powerful base is needed to carry out second dehydrohalogenation reaction *e.g.*, hot alcoholic KOH solution or alkoxide ion.

58. (a) : Alkynes react with strong bases like $NaNH_2$ to form sodium acetylide derivative known as acetylides.

$$H-C\equiv C-H + NaNH_2 \longrightarrow H-C\equiv\bar{C}\overset{+}{Na} + 1/2H_2$$

59. (d) : Shortest C–C distance (1.20 Å) is in acetylene.

60. (a) :

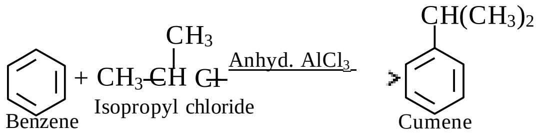


The formation of C – H bond in acetylene involves *sp*-hybridised carbon atom. Since *s*-electrons are closer to the nucleus than *p*-electrons, the electrons present in a bond having more *s*-character will be more closer to the nucleus. In alkynes *s* character is 50%, the electrons constituting this bond are more strongly bonded by the carbon nucleus. Thus, acetylenic C-atom becomes more electronegative in comparison to sp^2, sp^3 and hence the hydrogen atom present on carbon atom ($\equiv C$– H) can be easily removed.

61. (d) : All the three reagents except ammoniacal $AgNO_3$ reacts with 1, 2 and 4 compounds. The compound 3 possessing the terminal alkyne only reacts with ammoniacal $AgNO_3$ and thus can be distinguished from 1, 2 and 4 compounds.

62. (c) : The attacking species in the reaction given in option (c) is an electrophile *i.e.*, $\overset{\delta+}{Cl}$. Therefore, it is an electrophilic substitution reaction.

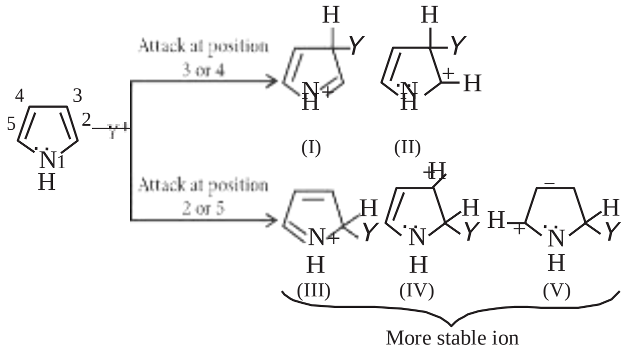
63. (d) : Friedel–Crafts reaction :



Chlorobenzene, bromobenzene and chloroethene are not suitable halide components as C–*X* bond acquires some double bond character due to resonance of lone pair of electrons with π-bond.

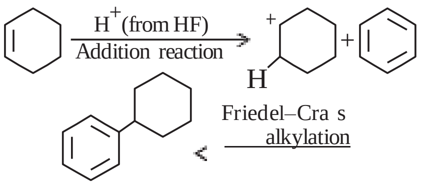
64. (a) : Biphenyl is coplanar as all C-atoms are sp^2 hybridised.

65. (d) : Pyrrole has maximum electron density on 2 and 5. It generally reacts with electrophiles at the C-2 or C-5 due to the highest degree of stability of the protonated intermediate.



Attack at position 3 or 4 yields a carbocation that is a hybrid of structures (I) and (II). Attack at position 2 or 5 yields a carbocation that is a hybrid not only of structures (III) and (IV) (analogous to I and II) but also of structure (V). The extra stabilization conferred by (V) makes this ion the more stable one. Also, attack at position 2 or 5 is faster because the developing positive charge is accommodated by three atoms of the ring instead of only two.

66. (c) :

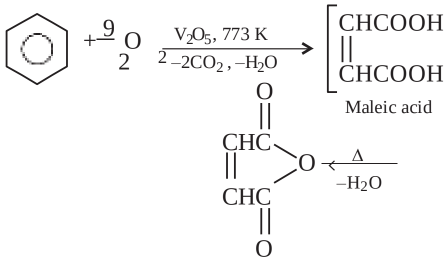


67. (d) : Mechanism of nitration is :

$HNO_3 + 2H_2SO_4 \rightarrow NO_2^+ + 2HSO_4^- + H_3O^+$

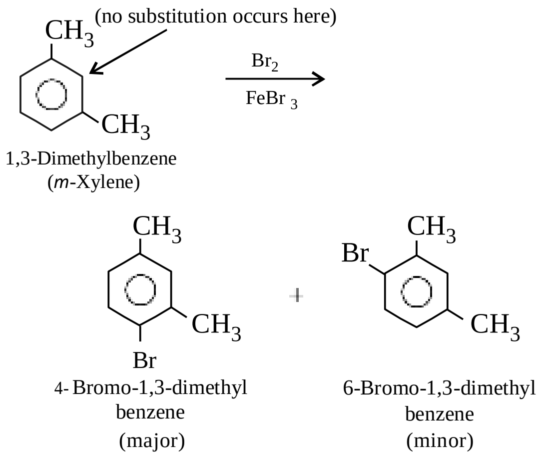
If a large amount of $KHSO_4$ is added then conc. of HSO_4^- ions increases and the reaction will be shifted in backward direction hence, the rate of nitration will be slower.

68. (a) :



Maleic anhydride

69. (c) : — CH_3 group is o,p-directing. Because of crowding, no substitution occurs at the carbon atom between the two — CH_3 groups in m-xylene, even though two — CH_3 groups activate that position.

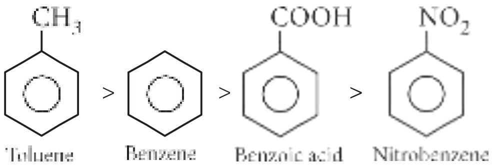


70. (b): —NO_2 is most deactivating due to $-I$ and $-M$ effect.

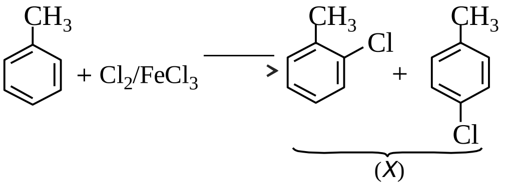
71. (a) : Nitrobenzene is strongly deactivated, hence will not undergo Friedel-Crafts reaction.

72. (d) : The molecules which do not satisfy Huckel rule or $(4n + 2)\pi$-electron rule are said to be non-aromatic. The compound (d) has total $4\pi e^-$. It does not follow $(4n + 2)\pi$ rule. So, it is non-aromatic compound.

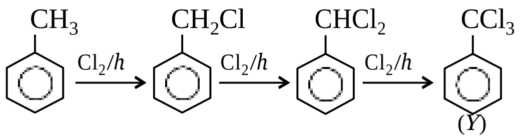
73. (c) : As the $+I$ effect increases reactivity towards electrophilic reactions increases and as $-I$ or $-M$ effect increases, reactivity towards electrophilic reactions decreases. Thus, the order is



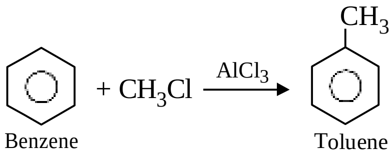
74. (c) : The reaction of Cl_2, in presence of $FeCl_3$, with toluene yields a ring substitution product.



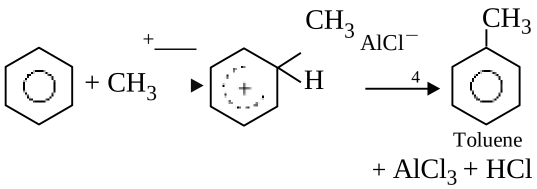
In presence of sunlight, free radical reaction takes place.



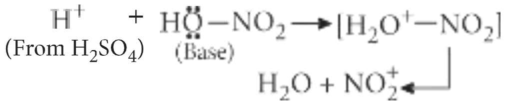
75. (d) : This is Friedel-Crafts alkylation.



Mechanism: $CH_3Cl + AlCl_3 \rightarrow \underset{\text{Lewis acid}}{AlCl_4^-} + \underset{\text{Electrophile}}{CH_3^+}$



76. (b) :



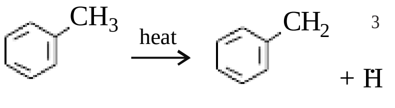
$$\underset{(\text{From } H_2SO_4)}{H^+} + \underset{(\text{Base})}{H\ddot{O}-NO_2} \rightarrow [H_2O^+-NO_2] \rightarrow H_2O + NO_2^+$$

77. (a) : Groups like, — Cl and — NO_2 shows –*I* effect. –*I* groups attached to the benzene ring decrease the electron density and hence less prone to electrophilic attack. — OH not only shows –*I* effect but also +*M* effect which predominates the –*I* character and electron density is increased in the benzene ring which facilitates electrophilic attack.

78. (c) : Benzene having any activating group *i.e.*, — OH, — *R* undergoes electrophilic substitution easily as compared to benzene itself. Thus, toluene and phenol undergo electrophilic substitution easily. Chlorine due to –*I*-effect deactivates the ring. So, it is difficult to carry out the electrophilic substitution in chlorobenzene. Hence, the order is $C_6H_5OH > C_6H_5CH_3 > C_6H_6 > C_6H_5Cl$.

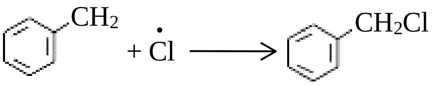
79. (c) : $C_6H_5H + H_2C == CH_2 \xrightarrow[95°C,\ \text{pressure}]{AlCl_3,\ HCl} C_6H_5CH_2CH_3$

80. (a) :

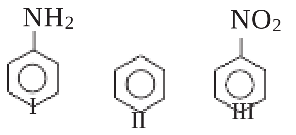


Benzyl free radical

$$Cl_2 \xrightarrow[\text{sunlight}]{\text{Energy}} 2\dot{C}l$$

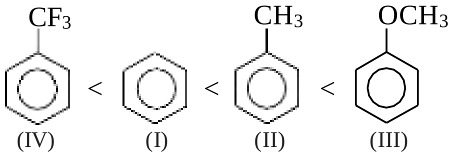


81. (d) :

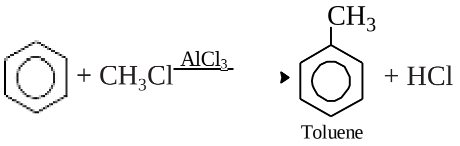


— NH_2 group is electron donating hence increases electron density on ring. Benzene is also electron rich due to delocalisation of electrons. — NO_2 group is electron withdrawing hence, decreases electron density on ring. Thus, correct order for electrophilic substitution is I > II > III.

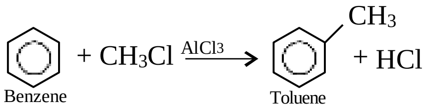
82. (a) : Due to –*I* effect of F atom, —CF_3 on benzene ring, deactivates the ring and does not favour electrophilic substitution. While — CH_3 and — OCH_3 are electron-donating group which favours electrophilic substitution in the benzene ring at '*ortho*' and '*para*' positions. The +*I* effect of — OCH_3 is more than — CH_3, therefore the correct order for electrophilic substitution is



83. (a) : In Friedel-Crafts reaction toluene is obtained by the action of CH_3Cl on benzene in presence of $AlCl_3$.



84. (a) : In Friedel-Crafts reaction, an alkyl group is introduced into the benzene ring in presence of a Lewis acid ($AlCl_3$) catalyst. The reaction is



85. (a) : — OH, — Cl and — CH_3 groups in benzene are *ortho, para* directing groups. But among these — OH group is strongly activating while — CH_3 is weakly activating and — Cl is deactivating. Thus, phenol will be most easily attacked by an electrophile.

86. (a) : Arenes undergo electrophilic substitution reactions and are resistant to addition reactions, due to delocalisation of π-electrons. These are also stabilized by resonance.

87. (a) : In structure III, withdrawal of electrons by —NO_2 causes decrease in reaction rate while in structure I, there is electron releasing effect by —OCH_3 group which accelerates the reaction.

The order of reactivity towards electrophile is :

I > II > III

88. (c) : $Cl_2 + FeCl_3 \longrightarrow FeCl_4^- + Cl^+$

89. (c) : Nitronium ion (NO_2^+) is an electrophile that actually attacks the benzene ring.

90. (d) : Benzene shows Kekule structures which are resonating structures and these structures are separated by a double headed arrow (↔).

91. (c) : Due to resonance all the C – C bonds in the benzene possess same nature and the resonating structures are obtained because of the delocalisation of π-electrons.