

3. Structure of Atom

1. The electronic configuration of Cu (29) ____.
2. Elliptical orbits are introduced by ____.
3. The sub shell of the orbital for $l = 1$ is ____.
4. ____ is a group of wavelength.
5. Splitting of spectral lines due to magnetic field is called ____.
6. Splitting of spectral lines in the presence of electric field is called ____.
7. Stationary orbits are introduced by ____.
8. The electronic configuration of Cr is ____.
9. The elements in which outermost orbital's are completely filled are called ____.
10. Short notation of electron configuration is ____.
11. Shape of s-orbital is ()
a) Spherical b) Dumbbell c) Square d) Double Dumbbell
12. Shape of p-orbital is ()
a) Double Dumbbell b) Dumbbell c) Spherical d) Circle
13. Shape of d-orbital is ()
a) Spherical b) Dumbbell c) Double Dumbbell d) Square
14. In VIBGYOR which color having higher wavelength or lower frequency _____
()
a) Red b) Violet c) Indigo d) Orange
15. Which of the following is not correct? ()
a) $2p^6$ b) $3s^1$ c) $2d^3$ d) $4f^{12}$
16. Stable elements ()
a) Alkali b) Alkali Metals c) Inert d) None
17. The sub shell of the orbital for $l = 1$ is ()
a) s b) p c) d d) f

18. Short rotation of electron configuration is ()

- a) nl^x b) nl^n c) ln^x d) xn^l

19. Splitting of spectral lines in the presence of electric field is called ()

- a) Zeeman Effect b) Stark Effect c) Dispersion d) Limen Effect

20. What is 'n' value for L shell ()

- a) 1 b) 2 c) 3 d) 4

Answers

1) $[(Ar)4s^13d^{10}]$

2) Sommerfeld

3) p

4) Spectrum

5) Zeeman Effect

6) Stark Effect

7) Neil's Bohr

8) $[Ar]4s^13d^5$

9) Inert Gas

10) nl^x

11) a

12) b

13) c

14) a

15) c

16) c

17) b

18) a

19) b

20) b