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## CLASS X CHEMISTRY

**Q1) Give an example of a reaction which takes place in the presence of sunlight.**

**Q2) On heating blue coloured powder of copper (II) nitrate in a boiling tube, copper oxide (black), oxygen and a brown gas X is formed.**

- (a) Write a balanced chemical equation of the reaction.**
- (b) Identify the brown gas evolved.**
- (c) Identify the type of reaction.**

**(d) What could be the pH range of aqueous solution of the gas? Q3)**

**Define the following terms:-**

- (i) Ionic equation**
- (ii) Skeletal equation**

**Q4) How can you study experimentally that water is made of hydrogen and oxygen combined together in the ratio of 2:1 by volume? Q5) What do you mean by the term rancidity? Write the different methods to prevent rancidity. Q6) Why does the colour of silver nitrate change when a coin of copper is dipped in it? Q7) Write the balanced chemical equation for the following reactions and identify the type of reaction in**

**each case.**

**(i) Magnesium reacts with Iron (III) oxide at high temperature to form molten magnesium and Aluminium oxide.**

**(iii) Ethanol is burnt in air to form carbon dioxide, water and releases heat.**

**Q8) Identify the oxidizing agent (oxidant) in the following reactions:**

**(a)  $\text{Pb}_3\text{O}_4 + 8\text{HCl} \longrightarrow 3\text{PbCl}_2 + \text{Cl}_2 + \text{H}_2\text{O}$**

**(b)  $2\text{Mg} + \text{O}_2 \longrightarrow 2\text{MgO}$**

**(c)  $\text{CuO} + \text{H}_2 \longrightarrow \text{Cu} + \text{H}_2\text{O}$  Q9) Which among the following are physical or chemical changes?**

- (a) Evaporation of petrol.**
- (b) Curdling of milk.**
- (c) Sublimation of solid ammonium chloride.**
- (d) Burning of LPG.**
- (e) Rusting of iron. Q10) Why do we store silver chloride in dark colour bottles?**